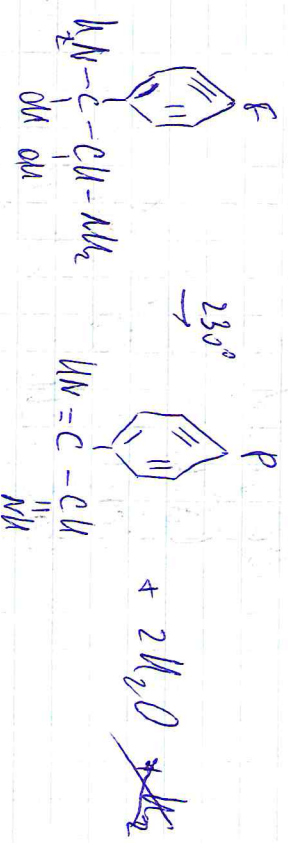
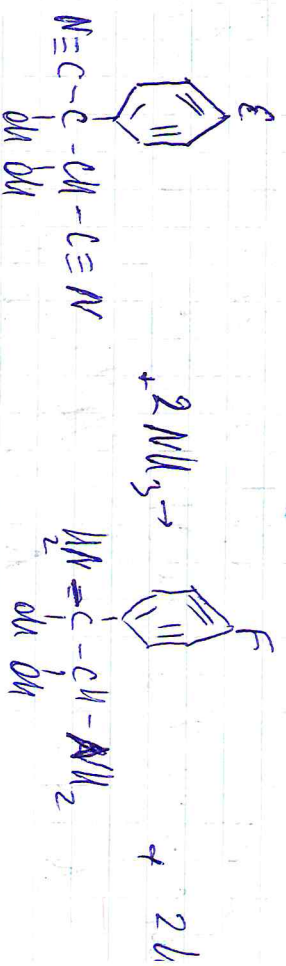


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20

1110

5-105  
 5-26  
 5-105  
 5-35  
 5-06  
 255



Problem:  $X - C_{10}H_{11}N_2$ ;  $P - C_8H_7N_2$  - Fluoride  
 11-4 Data:

$D_{12} \approx 4,655$   
 $Q \approx 126,4510^*$   
 $T \approx 400^\circ K$   
 $P \approx 59,77 kPa$

$N_2 + 3H_2 \xrightarrow{145^\circ} 2NH_3$   
 $H_2 + CuO \rightarrow Cu + H_2$   
 $N_2 + CuO \rightarrow N_2O + Cu$   
 $PV \approx \frac{P_0 V_0}{T_0}$

$$\frac{V_2 \cdot w_2}{w(C_{2H_2})} = \frac{V_2 \cdot w_2}{100\%}$$

$$V_2 = 0,247 \text{ л}$$

35

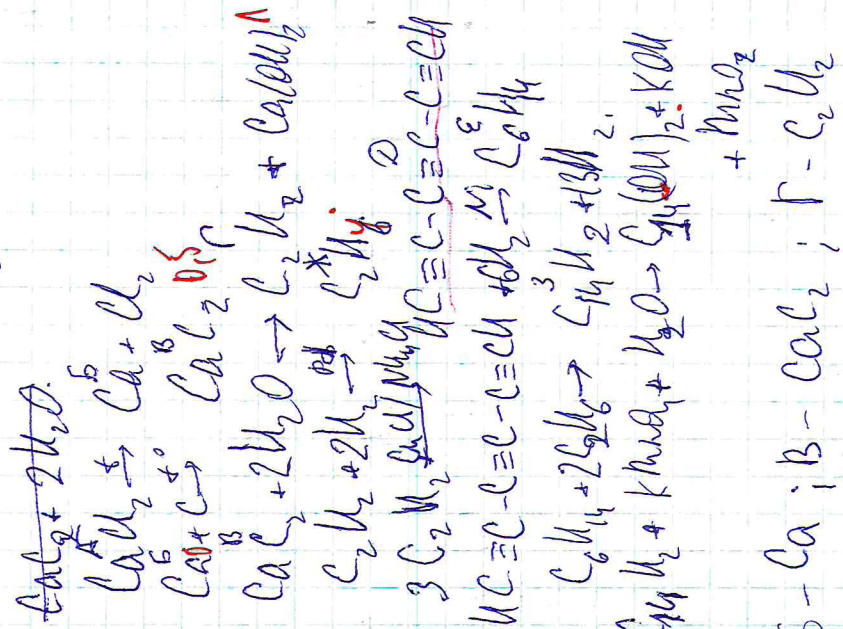
$$D \rightarrow C_2H_2; 2 - C_2H_4; 3 - C_2H_6; 4 - C_2H_2$$

Дано: 9011 моль

$$w(C) = 82,73\%$$

$$3 - C_2H_2$$

$$A - 3 - ?$$



35

$$Dано: A - CaC_2; B - Ca; C - C_2H_2$$

11-5-Дано:

$$t_1 = -18^\circ C$$

$$t_2 = -5^\circ C$$

$$t_3 = +20^\circ C$$

$$V_1 = 30 \text{ л}$$

$$V_2 = 6 \text{ л}$$

$$-18^\circ C - 30 \text{ л}$$

$$-5^\circ C - 6 \text{ л}$$

$$K = 0,22$$

$$+20^\circ C \cdot K = 2,2$$

$$20 \cdot 0,22 = 4,4$$

$$V_3 = ?$$

Реш: 14,4 л

0



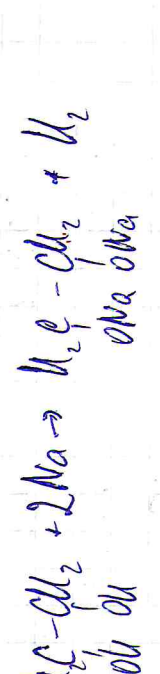
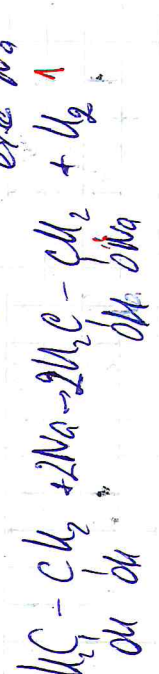
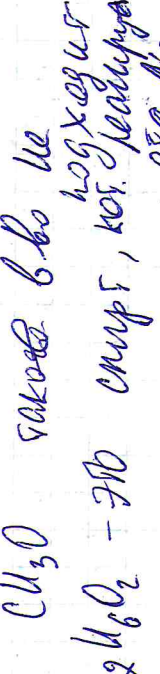


$n(C_6H_5O_2) = \frac{6.22}{92} \text{ моль}$   
 $n(C_6H_5O_2) = 0.0676 \text{ моль}$

$V(\text{воздух}) \approx 118 \text{ л}$   
 $n(H_2O) \approx 4.52$   
 $V(\text{воздух}) \approx 11.2 \text{ л}$   
 $O_2(\text{воздух}) \approx 14$

$n(C) = 2n(H_2O) = 9.04$   
 $n(H) = 2n(H_2O) = 9.04$   
 $n(O) = 2n(H_2O) = 9.04$   
 $n(C) : n(H) : n(O) = 1 : 1 : 1$

$12x + y + 16z = 31$   
 $15x + 16z = 31$   
 $x = 1, y = 3, z = 1$



Ответ: X - это смесь  $(C_6H_5O_2)$ ; при этом она дает реакцию Фелинга, в которой образуются и другие продукты.

Задача 11-1.  
 Дано: ...  
 Решение:

$n(C) = 2n(H_2O) = 9.04$   
 $n(H) = 2n(H_2O) = 9.04$   
 $n(O) = 2n(H_2O) = 9.04$   
 $n(C) : n(H) : n(O) = 1 : 1 : 1$

$n(C) = 2n(H_2O) = 9.04$   
 $n(H) = 2n(H_2O) = 9.04$   
 $n(O) = 2n(H_2O) = 9.04$   
 $n(C) : n(H) : n(O) = 1 : 1 : 1$

$X - ?$   
 $P - ?$   
 $Z - ?$

$n(C) = 2n(H_2O) = 9.04$   
 $n(H) = 2n(H_2O) = 9.04$   
 $n(O) = 2n(H_2O) = 9.04$   
 $n(C) : n(H) : n(O) = 1 : 1 : 1$

$n(C) = 2n(H_2O) = 9.04$   
 $n(H) = 2n(H_2O) = 9.04$   
 $n(O) = 2n(H_2O) = 9.04$   
 $n(C) : n(H) : n(O) = 1 : 1 : 1$

$n(C) = 2n(H_2O) = 9.04$   
 $n(H) = 2n(H_2O) = 9.04$   
 $n(O) = 2n(H_2O) = 9.04$   
 $n(C) : n(H) : n(O) = 1 : 1 : 1$

$n(C) = 2n(H_2O) = 9.04$   
 $n(H) = 2n(H_2O) = 9.04$   
 $n(O) = 2n(H_2O) = 9.04$   
 $n(C) : n(H) : n(O) = 1 : 1 : 1$

