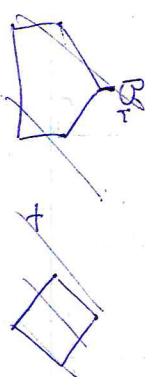
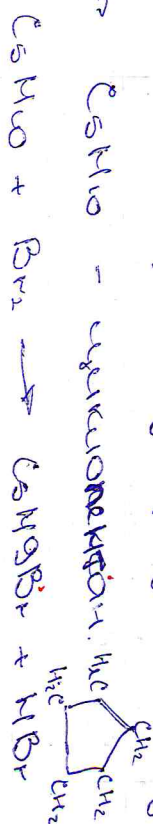


$$\begin{array}{r} 1105 \\ \text{I} - 105 \\ \text{II} - 15 \\ \text{III} - 25 \\ \text{IV} - 15 \\ \hline \text{V} - 0 \\ \hline 145 \end{array}$$
$$\begin{array}{l} \text{X} + \text{O}_2 \longrightarrow \text{XO}_2 + \text{H}_2\text{O} + \\ n = \Delta \text{ work} \quad \Delta H_m = 22.4 \text{ J/work} \end{array}$$

$$n(\text{CO}_2) = \frac{V}{V_m} = \frac{119.1}{22.4} = 5.316 \Rightarrow$$

[illegible]

↳  $\text{C}_5\text{H}_{10}$  - alkene ~~alkane~~.  $\text{H}_2\text{C}=\text{CH}-\text{CH}_2-\text{CH}_2-\text{CH}_3$

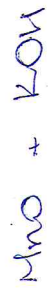


3- yuchogopog ~~233~~ leegban  
no opopuyga unnen buq exhy  
yg wacoba opus yuopoga p  
0.8475  $\Rightarrow$

$$= \frac{12x}{2x+2} \quad 0.875 \quad 1.472x = 0.875$$

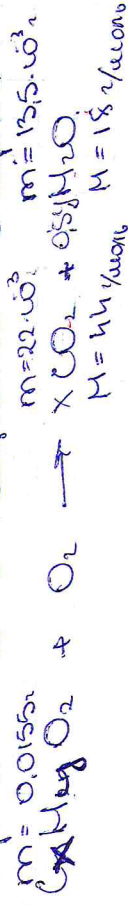
$$x \cdot y = 0,87 \cdot 13 = 11,724 \approx 12 \quad ; \quad 1,578 = 3:2 \Rightarrow C_6H_{10} \quad (\text{непредельный угле-})$$

2. E + \* - реакция Вюрца.



(11-3)

Органическое вещество - сульфат.



$n(CO_2) = 0,5 \cdot 10^{-3} \text{ моль}$   
 $n(H_2O) = 0,75 \cdot 10^{-3} \text{ моль}$

$x : y = 1 : 3$



$V = 202 \text{ см}^3 = 202 \cdot 10^{-6} \text{ м}^3$   
 $T = 200^\circ C = 473^\circ K$   
 $p = 1 \text{ атм} = 10^5 \text{ Па}$

$pV = \nu RT \Rightarrow \nu = \frac{pV}{RT} = \frac{202 \cdot 10^{-6} \cdot 10^5}{8,31 \cdot 473} =$

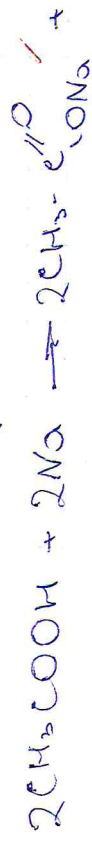
$= 0,00768 \text{ моль}$

$M(C_nH_{2n}O_2) = \frac{m}{\nu} = \frac{0,4202}{0,00768} \approx 64 \text{ г/моль}$

$15n = 64 - 32 \Rightarrow n \approx 2 \Rightarrow C_2H_4O_2$

Скорее всего 2 молекулы уксусной кислоты

Мономерный сульфат  $CH_3COOH$ .



(11-4)

$T = 400 K$ ;  $p = 99770 \text{ Па}$ ;  $V = 0,1$   
 $R = 8,31$

$\nu = \frac{pV}{RT} = \frac{99770 \cdot 0,1}{8,31 \cdot 400} = 3 \text{ моль}$

(конечное кол. вещества)